

Neutralizações

@quimicadaluanana
Parte I

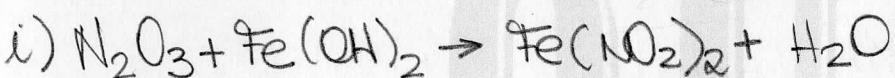
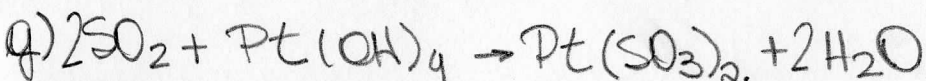
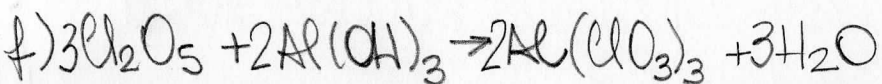
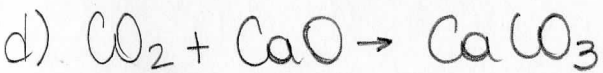
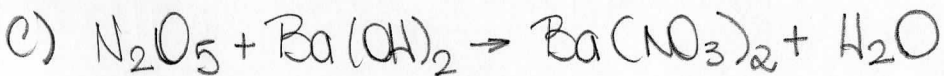
1)

- a) $\text{HF} + \text{Ca}(\text{OH})_2 \rightarrow \text{CaOHF} + \text{H}_2\text{O}$
- b) $\text{H}_2\text{SO}_4 + \text{KOH} \rightarrow \text{KHSO}_4 + \text{H}_2\text{O}$
- c) $\text{HNO}_3 + \text{Ba}(\text{OH})_2 \rightarrow \text{BaOHNO}_3 + \text{H}_2\text{O}$
- d) $\text{H}_3\text{PO}_4 + \text{Ca}(\text{OH})_2 \rightarrow \text{CaHPO}_4 + 2\text{H}_2\text{O}$
- e) $\text{HCl} + \text{Ca}(\text{OH})_2 \rightarrow \text{CaOHCl} + \text{H}_2\text{O}$
- f) $\text{HClO}_3 + \text{Al}(\text{OH})_3 \rightarrow \text{Al}(\text{OH})_2\text{ClO}_3 + \text{H}_2\text{O}$
- g) $\text{H}_2\text{S} + \text{Pt}(\text{OH})_4 \rightarrow \text{Pt}(\text{OH})_2\text{S} + 2\text{H}_2\text{O}$
- h) $\text{H}_2\text{CO}_3 + 3\text{NaOH} \rightarrow \text{Na}_3\text{OHCO}_3 + 2\text{H}_2\text{O}$
- i) $\text{HNO}_2 + \text{Fe}(\text{OH})_2 \rightarrow \text{FeOHNO}_2 + \text{H}_2\text{O}$
- j) $2\text{HF} + \text{Al}(\text{OH})_3 \rightarrow \text{Al}_2\text{OH}_2\text{F}_2 + 2\text{H}_2\text{O}$
- k) $\text{HIO}_3 + \text{Mg}(\text{OH})_2 \rightarrow \text{MgOHIO}_3 + \text{H}_2\text{O}$

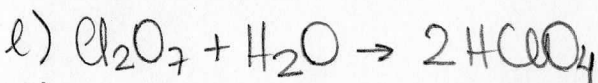
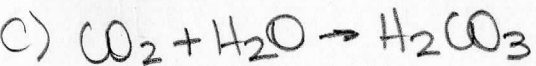
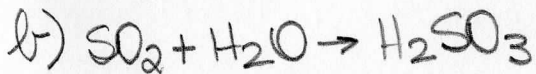
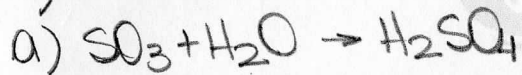
2)

- a) $2\text{HF} + \text{Ca}(\text{OH})_2 \rightarrow \text{CaF}_2 + 2\text{H}_2\text{O}$
- b) $\text{H}_2\text{SO}_4 + 2\text{KOH} \rightarrow \text{K}_2\text{SO}_4 + 2\text{H}_2\text{O}$
- c) $2\text{HNO}_3 + \text{Ba}(\text{OH})_2 \rightarrow \text{Ba}(\text{NO}_3)_2 + 2\text{H}_2\text{O}$
- d) $2\text{H}_3\text{PO}_4 + 3\text{Ca}(\text{OH})_2 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + 6\text{H}_2\text{O}$
- e) $2\text{HCl} + \text{Ca}(\text{OH})_2 \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$
- f) $3\text{HClO}_3 + \text{Al}(\text{OH})_3 \rightarrow \text{Al}(\text{ClO}_3)_3 + 3\text{H}_2\text{O}$
- g) $2\text{H}_2\text{S} + \text{Pt}(\text{OH})_4 \rightarrow \text{PtS}_2 + 4\text{H}_2\text{O}$
- h) $\text{H}_2\text{CO}_3 + 2\text{NaOH} \rightarrow \text{Na}_2\text{CO}_3 + 2\text{H}_2\text{O}$
- i) $2\text{HNO}_2 + 2\text{Fe}(\text{OH})_2 \rightarrow \text{Fe}(\text{NO}_2)_2 + 2\text{H}_2\text{O}$
- j) $3\text{HF} + \text{Al}(\text{OH})_3 \rightarrow \text{AlF}_3 + 3\text{H}_2\text{O}$
- k) $2\text{HIO}_3 + \text{Mg}(\text{OH})_2 \rightarrow \text{Mg}(\text{IO}_3)_2 + 2\text{H}_2\text{O}$

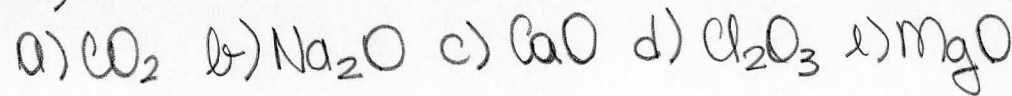
3)



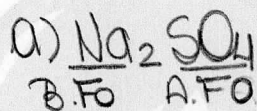
4)



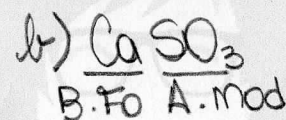
5)



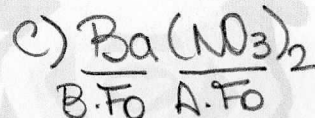
6)



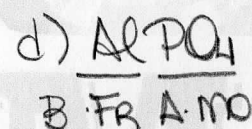
sulfato de sódio; solúvel, pH = 7
s/ hidrólise



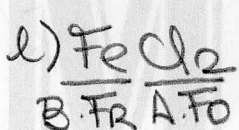
sulfite de cálcio, insolúvel, pH > 7



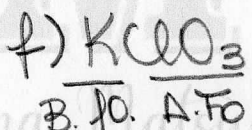
Nitrato de Bário, solúvel, pH = 7
s/ hidrólise



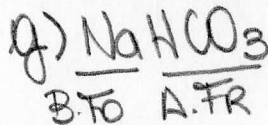
fosfato de alumínio, insolúvel, pH < 7



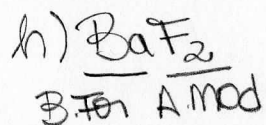
cloreto de ferro(II); solúvel, pH < 7
cloreto ferroso



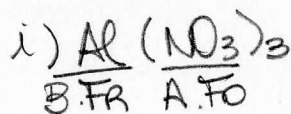
clorato de potássio, solúvel, pH = 7
s/ hidrólise



Bicarbonato de sódio
hidrogenocarbonato de sódio, solúvel, pH > 7
carbonato ácido de sódio



fluoreto de Bário, insolúvel, pH > 7



Nitrato de Alumínio, solúvel, pH < 7

6)

J) $\underline{K_3PO_4}$ fosfato de potássio
B.Fo A.FR solúvel, pH > 7

K) $\underline{Na_3PO_4}$ fosfato de sódio
B.Fo A.FR solúvel, pH > 7



QUÍMICA

Luana Matsunaga

7)

a) óxido de cálcio, óx. Básica

b) trióxido de enxofre, óx. Ácido

c) óxido de alumínio, óx. anfótero

d) peróxido de hidrogênio

e) óxido de ferro (II); óx. Básica
óxido ferroso

f) óxido de cobre (I); óx. Básica
óxido cuproso

g) pentóxido de dinitrogênio, óx. Ácido

h) monóxido de nitrogênio, óx. neutro

i) óxido de zinco, óx. anfótero